Topic Reminder Q14
Weak and Strong Acids

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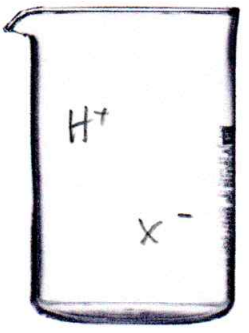
1. Calculate the pH of a 0.15M HCl.

$$-\log(0.15) = 0.82$$



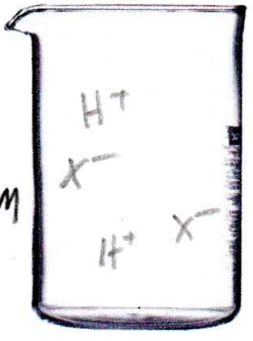
2. Draw a picture of the solution in the beaker to the right.

3. Draw a picture of a 0.3M solution of HCl in the beaker to the right.

$$-\log(0.3) = 0.52$$

4. Calculate the pH of a 0.15M HF solution (Ka = 6.4E-4)

HF $\qquad 6.4E-4 = \frac{x^2}{.15} \qquad x = 0.0097M$



5. Draw a picture of a 0.15M HF solution in the beaker to the right.

6. Write the hydrolysis reaction for HF

$$HF + H_2O \rightleftharpoons H_3O^+ + F^-$$
A B C A C B

7. Label the Acid base conjugate pairs of #6.